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= 12.00 mmol - 5.40 mmol Unreacted ethanoic acid = 6.60 mmol Since 18.00 mL of NaOH(aq) was added to 20.00 mL of ethanoic acid solution, the total volume is now 38.00 mL. Solution: The concentration of unreacted ethanoic acid, [CH₃COOH(aq)], can be determined:

$$\frac{6.60 \text{ mmol}}{38.00 \text{ mL}} \times 1000 \text{ mL} = 0.174 \text{ mol/L}$$

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$\text{HC}_2\text{H}_3\text{O}_2(\text{aq})$ 23.2 [$\text{HC}_2\text{H}_3\text{O}_2(\text{aq})$] (6.60
mmol) = 38.00 mL

Section 8.7: Acid-Base Titration **Tutorial 1 Practice, page 547**

2 Chemistry, First Canadian Edition 1.17

An organic compound consists of molecules made from carbon, hydrogen and oxygen atoms. There can be other

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atoms – for example, nitrogen, sulfur.

1.19 1. Acetylsalicylic acid, the common mild pain reliever, is a derivative of the natural product, salicin, obtained from willow bark. 2.

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Unit 1 Organic Chemistry - Mr. Arthur's Science Page

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