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The Lower Electronegativity transfers to the higher electronegativity) t. Ionic Bond. Ions pack in a crystal lattice structure. Lattice Energy. Cation {+} and anion {-} create _____

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(nonmetals) ... Modern Chemistry Chapter 6 Vocabulary. 20 terms.

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Modern Chemistry 45 Chemical Bonding CHAPTER 6 REVIEW
Chemical Bonding SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. ____ The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. ____ In a crystal of an ionic compound, each cation is surrounded by a ...

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CHAPTER 6 REVIEW Chemical Bonding SECTION 2 SHORT ANSWER Answer the following questions in the space provided.

1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the

6 Chemical Bonding

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CHAPTER 6 REVIEW Chemical Bonding Modern Chemistry 45
Chemical Bonding CHAPTER 6 REVIEW Chemical Bonding
SECTION 3 SHORT ANSWER Answer the following questions in the space provided
1 _____ The notation for sodium chloride, NaCl, stands for one (a) formula unit (c) crystal (b) molecule (d) atom
2 _____ In a crystal of an ionic compound, each ...

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Chemical Bonding CHAPTER 6 Chemical Bonding SECTION 1
Introduction to Chemical Bonding OBJECTIVES 1. Define Chemical bond. 2. Explain why most atoms form chemical bonds. 3. Describe ionic and covalent bonding.. 4.

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CHAPTER 12 REVIEW Solutions SECTION 2 SHORT ANSWER
Answer the following questions in the space provided. 1. The

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following are statements about the dissolving process. Explain each one at the molecular level. a. Increasing the pressure of a solute gas above a liquid solution increases the solubility of the gas

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